Types of Chemical Reactions

Learning Goals: I will be able to ...

- identify and define the main types of chemical reactions
- predict the products of synthesis and decomposition reactions
- predict the products of *single displacement* and *double displacement* reactions
- predict the products of neutralization reactions
- write balanced chemical equations for these reactions

Synthesis Reactions

- →Two or more reactants combine to form one new product.
- →Represented by:
- →Example: A + B -> AB

$$2 H_2(g) + O_2(g) -> 2H_2O(I)$$

Many greenhouse gases are formed through a synthesis reaction.

For example;

$$N_2(g) + O_2(g) \rightarrow 2NO(g)$$

So what do all these (g), (l), (aq), and (s)'s mean?????

□Abbreviations for the States of Reactants and Products

State	Abbreviation	Examples(at room temp)
Solid	(s)	Most metals, ionic compounds such as NaCl
Liquid	(I)	H2O, H2O2, Hg
Gas	(g)	H2, O2, F2, N2, Cl2, SO2
Aqueous Solution	(aq)	Any soluble solid such as NaCl

Decomposition Reactions

- →A compound is broken down into two or more smaller products.
- \rightarrow Many of these reactions require heat (Δ) to occur. Some require an electric current to occur.
- →Represented by:

AB -> A + B



2AgCl (s) -> 2Ag (s) +
$$Cl_2$$
 (g)

Single Displacement Reactions

- →One element replaces another element in a compound (usually an ionic compound).
- →Represented by:

$$A + BC (aq) -> B + AC (aq)$$
 $A + BC (aq) -> C + BA (aq)$

For example:

Lithium + Barium Nitride Barium + Lithium Nitride Li +
$$Ba_3N_2$$
 -> $Ba + Li_3N$

Double Displacement Reactions

- →The cation of one compound exchanges place with the cation of another compound (usually ionic compounds).
- →Represented by:

- →A <u>true</u> double displacement reaction occurs when one of the products is:
 - 1. a precipitate (s) 2. a gas (g) 3.water (l)

5. Combustions Reactions

A combustion reaction occurs when a hydrocarbon (organic molecule) reacts with oxygen to produce energy.

□Combustion reactions are usually in the form:

$$CnH_2n + O_2(g) \longrightarrow H_2O(I) + CO_2(g) + Energy$$

□For Example:

$$CH_4(g) + O_2(g) \rightarrow 2H_2O(g) + CO_{2(}g) + Energy$$

This is the Complete Combustion of Methane gas

Fuel + Oxygen -> Water and Carbon Dioxide (hydro carbon)

Try These:

Identify the following reactions as one of the given types:

$$2AgF + NiCl_2 \rightarrow 2AgCl + NiF_2$$

$$2K + MgBr_2 \rightarrow 2 KBr + Mg$$

Li +
$$Br_2 \rightarrow 2LiBr$$

$$Na_2CO_3 \rightarrow CO_2 + Na_2O$$

$$C_2H_8 + 4O_2 \rightarrow 4H_2O + 2CO_2$$

